

# Water and living organisms

1. The alkaline properties of solution with a pH  $>7$  are due to:
  - hydroxide ions
  - hydroxonium ions
  - oxide ions
  - hydrogen ions
2. The acidic properties of solution with a pH  $<7$  are due to:
  - hydroxide ions
  - oxide ions
  - hydroxonium ions
  - hydrogen ions
3. In a water molecule, each bond between the oxygen atom and the two hydrogen atoms are best described as:
  - a hydrogen bond
  - a single polar covalent bond, with a small positive charge on the hydrogens and a small negative charge on the oxygen
  - a single non-polar covalent bond, with a small negative charge on the hydrogens and a small positive charge on the oxygen
  - a single non-polar covalent bond